

Experiment 39

FV 6/4/12

MAGNETIC BEHAVIOR AND ELECTRON CONFIGURATIONS

MATERIALS: Analytical balance, strong neodymium magnet, samples of sodium chloride, potassium chloride, magnesium oxide, calcium oxide, magnesium chloride, calcium chloride, aluminum chloride, aluminum oxide, iron(III) chloride, cobalt(II) chloride, copper(II) oxide, ammonium dichromate, water, ethanol, sulfur and iodine.

PURPOSE: The purpose of this experiment is to determine the electron configurations of ions in a group of ionic compounds by investigating the behavior of the compounds when placed into a magnetic field.

LEARNING OBJECTIVES: By the end of this experiment, the student should demonstrate the following proficiencies:

1. Determine the magnetic properties of several materials.
2. Relate the magnetic behavior of an ionic substance to the electron configurations of the component ions.

DISCUSSION:

Electron configurations are extremely important in determining many chemical and physical properties of elements and compounds. The electron configuration of an atom or ion helps to determine such properties as:

1. Valence - chemical combining power
2. Metallic or non-metallic behavior
3. Atomic and ionic radius
4. Melting and boiling point
5. Polarizability
6. Magnetic properties
7. Color

Magnetism - The behavior of a substance when exposed to a magnetic field is related to the existence of paired and unpaired electrons in the material. We will consider two types of magnetism in this experiment.

Diamagnetism - When an external magnetic field is brought close to a diamagnetic material, a very weak magnetic field is created which opposes the applied magnetic field. The magnet and material repel each other. All electrons contribute to the property of diamagnetism but in order for a material to be diamagnetic, all of the electrons must be paired.

Paramagnetism - When an external magnetic field is brought close to a paramagnetic material, the magnet and material are attracted to each other. In order to be paramagnetic, there must be at least one unpaired electron in the material. Paramagnetism is weak but quite a bit stronger than diamagnetism. So, in a material that contains both paired and unpaired electrons, the paramagnetism is generally the stronger effect and the net result is an attraction of the material and the magnet.

Magnetic properties of a material can be examined by noting the behavior of a magnet on an analytical balance when a material is brought close. If the material is placed near the top of the magnet, a diamagnetic material will repel the magnet. The magnet will appear to be heavier, noted by an increase in the mass value. A paramagnetic material will attract the magnet so that the magnet will appear to be lighter, noted by a decrease in the mass value. The paramagnetic effect is considerably stronger. In this lab you will determine the magnetic behavior of certain classes of compounds or elements. You will then be asked to state how the magnetic behavior correlates with the electron configurations of the ions/atoms in the substance.

PROCEDURE:

1. Place a magnet on an analytical balance. Record the mass.
2. Position a container holding the sample near the top of the magnet. Record the new mass.

If the magnet appears lighter in the presence of the sample, the sample is attracting the magnet and is paramagnetic (at least one electron is unpaired).

If the magnet appears heavier in the presence of the sample, the sample is repelling the magnet and is diamagnetic (all electrons are paired).

NOTES:

1. The balances are very sensitive to any movement. Ignore fluctuations of the balance that occur in the ten thousandths place. To minimize the turbulence, keep the side doors closed and insert the samples through the top door of the balance. Lower the samples slowly. Do not let the sample contact the magnet.
2. The containers and our hands also display these magnetic properties.

How diamagnetic are the containers? To get a baseline, do a run with the empty container.

Baseline	mass (g) magnet alone	mass (g) magnet +container	Δ mass (g) (container - magnet)	Magnetism (P or D)
Container				

How diamagnetic are you? To test your own diamagnetism, just place your hand close to the magnet.

Baseline	mass (g) magnet alone	mass (g) magnet +hand	Δ mass (g) (hand - magnet)	Magnetism (P or D)
Your hand				

Name: _____

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EXPT. 39 - DATA COLLECTION AND ANALYSIS**Part A: Simple 1:1 Ionic Compounds**

Ionic compounds are formed by combining positively charged metal ions with negatively charged non-metal ions. We will examine at some possible electron configurations for the ions in some simple ionic compounds. Assume the metal ions can have a charge as high as +3, and the non-metal ions can have a charge as large as -3.

The compounds below have a 1:1 stoichiometry. Determine whether each compound is paramagnetic or diamagnetic.

Compound	mass (g) magnet alone	mass (g) magnet + compound	Δ mass (g) (compound - magnet)	Magnetism (P or D)
A1 sodium chloride				
A2 potassium chloride				
A3 calcium oxide				
A4 magnesium oxide				

Part B: Ionic Compounds without 1:1 Stoichiometry

Samples of magnesium chloride, calcium chloride, aluminum chloride and aluminum oxide are available for experimentation. Use the same procedure as above to determine the magnetic properties of these ionic compounds and their chemical formulas.

Compound	mass (g) magnet alone	mass (g) magnet + compound	Δ mass (g) (compound - magnet)	Magnetism (P or D)
B1 magnesium chloride				
B2 calcium chloride				
B3 aluminum chloride				
B4 aluminum oxide				

Part C: Ionic Compounds with Transition Metals

In this section, the charges of the metal ions are known.

Test the magnetic properties of each of the compounds listed above.

Compound	mass (g) magnet alone	mass (g) magnet + compound	Δ mass (g) (compound - magnet)	Magnetism (P or D)
C1 iron(III) chloride				
C2 cobalt(II) chloride				
C3 copper(II) oxide				
C4 ammonium dichromate				

Part D: Molecular Substances:

Molecular compounds differ from ionic compounds in that they form discrete, independent units that we call molecules. The attraction of positive and negative ions goes a long way to explain why ionic compounds form. The bonding in molecular compounds is a bit harder to understand. Examine the magnetic properties of the following molecular substances: water, ethanol, iodine, and sulfur.

Compound/Element	mass (g) magnet alone	mass (g) magnet + compound	Δ mass (g) (compound - magnet)	Magnetism (P or D)
D1 H ₂ O (l)				
D2 I ₂ (s)				
D3 C ₂ H ₅ OH (l)				
D4 S ₈ (s)				

QUESTIONS:

1. For the samples in Part A, use the possible electron configurations for each ion, as determined in your Pre-Lab, to give the chemical formulas that are consistent with your experimental results.

A1 sodium chloride :
A2 potassium chloride :
A3 calcium oxide :
A4 magnesium oxide :

Are your metal charges consistent with what you have already learned about common ion charges? Explain.

2. In Part B, use possible electron configurations of the ions to give the chemical formulas that are consistent with your experimental results.

B1 magnesium chloride :
B2 calcium chloride :
B3 aluminum chloride :
B4 aluminum oxide :

3. The compounds above have more than one possible empirical formula based on their magnetic behavior. To determine the actual empirical formula, another experiment must be performed. Suppose a 2.15 g sample of magnesium chloride is decomposed using an electrolytic cell to drive off the chlorine as Cl_2 gas. If the remaining metallic magnesium has a mass of 0.55 g, what is the empirical formula of the original magnesium chloride sample?

4. In Part C, for paramagnetic compounds involving transition metals, which electrons are responsible for the paramagnetic behavior?

5. Based upon your results in Part D, does bonding in molecular compounds tend to lead to pairing or unpairing of electrons?

Name _____

Date _____

Section _____

EXPT. 39 PRE-LAB: For each of the atoms/ions, complete the following table. P = paramagnetic D = diamagnetic

Atom/Ion	Electron Configuration (short-hand notation is OK)	# of unpaired electrons	Magnetism (P or D)
Na			
Na ⁺			
Na ²⁺			
K			
K ⁺			
K ²⁺			
Mg			
Mg ⁺			
Mg ²⁺			
Ca			
Ca ⁺			
Ca ²⁺			
Al			
Al ⁺			
Al ²⁺			
Al ³⁺			
Fe			
Fe ³⁺			
Co			
Co ²⁺			
Cr			
Cr ⁶⁺			
Cu			
Cu ²⁺			
Cl			
Cl ⁻			
Cl ²⁻			
Cl ³⁻			
O			
O ⁻			
O ²⁻			
O ³⁻			